



Handwritten lines and numbers, possibly a ledger or account entry.

Handwritten calculations:
$$\begin{array}{r} 4450 \\ 1150 \\ \hline \end{array}$$
 Below the calculation, there is a signature and the date '19.12.61'.

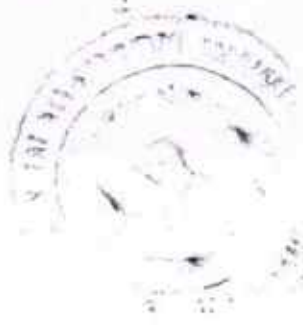
Vertical columns of handwritten text in Hindi, likely a list of items or a detailed account.

A large section of handwritten text in Hindi, appearing to be a detailed ledger or account with multiple columns and entries.

Handwritten text at the bottom of the page, possibly a signature or a concluding note.

BHO

26/12/62



Enumerate as under

- 1) Sairashtra Ch. Mungji
- 2) Ramsh Ch. Mungji S/O D/- Ramsh Ch. Mungji of Rahata P. Khudk of 24 Op of 1st kind of 1st service.
- 3) Nani Bala Bhang W/O D/- Jalindra Mahesh Bhang of Rahata P. Khudk of 1st kind of 1st service as Const. India along for (a) Rahata Bhang (b) Sairashtra Mahesh Bhang (c) Karabanda Mahesh Bhang (d) Badamsh Khudk

12-4-57
19th
Dec 62
Sairashtra Ch. Mungji

Nani Bala Bhang
D/- Jalindra Mahesh Bhang
Rahata
Khudk
Head
Shirani ke for set
and as guide in this unit sum
(i) Parthe Sarathi Bhang (ii) Sairashtra Bhang (iii) Karabanda Bhang

7123
19.12.62

Ranjit Rayin Paul
Munendra Nath Paul
Rahata
Khudk
Head
Sairashtra

7124



7125

19.12.62

Handwritten signature or scribble at the bottom left.

1. Yield - the amount of product obtained from a reaction
 2. Reaction rate - the speed at which a reaction takes place
 3. Equilibrium - the state where the concentrations of reactants and products are constant
 4. Activation energy - the minimum energy required for a reaction to occur
 5. Enthalpy change - the heat energy change during a reaction
 6. Entropy - a measure of the disorder or randomness of a system
 7. Free energy - the energy available to do work in a system
 8. Rate constant - a proportionality constant in the rate equation
 9. Order of reaction - the sum of the powers to which the concentrations of the reactants are raised in the rate equation
 10. Half-life - the time taken for the concentration of a reactant to decrease to half its initial value
 11. Collision theory - a model for reaction rates based on the collision of particles
 12. Mechanism - a series of steps that describe how a reaction takes place
 13. Transition state - the high energy state of a reaction intermediate
 14. Rate-determining step - the slowest step in a reaction mechanism
 15. Stoichiometry - the study of the quantitative relationships between the reactants and products in a chemical reaction
 16. Molar mass - the mass of one mole of a substance
 17. Empirical formula - the simplest whole number ratio of atoms in a compound
 18. Molecular formula - the actual number of atoms of each element in a molecule
 19. Structural formula - a diagram showing the arrangement of atoms in a molecule
 20. Isomerism - the existence of molecules with the same molecular formula but different structural formulas

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Handwritten title at the top of the page.

Handwritten notes in the top right section, including phrases like "The main problem is..."

Handwritten notes in the middle right section, starting with "The history of..."

Handwritten notes in the middle left section, starting with "The main problem is..."

Handwritten title for the bottom section.

Handwritten notes in the bottom section, starting with "The main problem is..."

